Student Performance by Question **Chemistry** (Multiple-Choice ONLY) Teacher: _____

Analysis of Performance by Question of each student targeted for Remediation or Intervention

Enter a "1" in each box to indicate a missed question.

	Scientific Investigation										Atomic Structure and Periodic Relationships								Chemical Formu						
Student Name	Calculate results of a controlled experiment using significant digits conventions.	Recognize and follow safe and appropriate laboratory procedures.	Recognize and follow safe and appropriate laboratory procedures.	Analyze results of a controlled experiment using significant digits conventions.	Identify the correct equipment to collect the results of a safe controlled experiment.	Use procedural analysis to defend the conclusions of a controlled investigation.	Calculate the percent error in the results of a controlled experiment.	Interpret results using appropriate error analysis.	Recognize and follow safe and appropriate laboratory procedures.	Interpret results using appropriate error analysis.	Recognize that periodic trends can be determined from the periodic table.	Apply information provided by the periodic table to construct models.	Identify characteristics of subatomic particles.	Apply information provided by the periodic table to construct models.	Apply information provided by the periodic table to perform calculations offormula mass.	Recognize that elements are unique.	Use bonding principles to explain the interaction of substances during chemical changes.	Apply information provided by the periodic table to make comparisons of matter.	Classify chemical equations.	Describe compounds and bonding to provide appropriate chemical formulas.	Describe compounds and bonding to provide appropriate chemical names.	Interpret factors that affect equilibrium and kinetics.	Describe compounds and bonding to provide appropriate chemical names.	Interpret factors that affect equilibrium and kinetics.	Balance chemical equations.

Total	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0	0

ulas and Reactions									Molar Relationships								Phases of Matter and Kinetic Molecular Theory								
Use bonding principles to explain the interaction of substances during chemical changes.	Apply information provided by the periodic table to construct models.	Identify the empirical formula of a substance.	Balance chemical equations.	Classify chemical equations.	Recognize that properties can be determined from the periodic table.	Classify chemical equations.	Balance chemical equations.	Interpret factors that affect equilibrium and kinetics.	Apply molar relationships to perform calculations involving molar concentrations.	Apply molar relationships to perform calculations involving molar concentrations.	Apply molar relationships to perform calculations involving stoichiometry.	Interpret factors that affect solution chemistry.	Apply molar relationships to perform calculations involving stoichiometry.	Apply molar relationships to perform calculations involving molar concentrations.	Interpret factors that affect solution chemistry.	Apply molar relationships to perform calculations involving molzf concentrations.	Apply the Kinetic Molecular Theory to predict the behavior of matter.	Recognize that the Kinetic Molecular Theory explains the behavior of matter.	Apply the Kinetic Molecular Theory to calculations involving specific heat.	Infer the behavior of matter based on the Kinetic Molecular Theory.	Apply molar relationships to perform calculations involving molar conversions.	Apply the Kinetic Molecular Theory to predict the behavior of matter.	Recognize that the Kinetic Molecular Theory explains the behaviorof matter.	Interpret factors that affect equilibrium and kinetics.	Total
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